

Note: - You have four choices for each objective type question as A, B, C and D. The choice which you think is correct; fill that circle in front of that question number in your answer book. Use marker or pen to fill the circles. Cutting or filling up two or more circles will result no mark.

SECTION-A

Q.1	Questions	A	B	C	D
1.	A solution of glucose is 10% w/v. The volume in which 1g mole of its dissolved will be:	1 dm ³	1.8 dm ³	200 cm ³	900 cm ³
2.	The pH of 10 ⁻³ mole dm ⁻³ of an aqueous solution of H ₂ SO ₄ is:	3.0	2.7	2.0	1.5
3.	Calorie is equivalent to:	0.4184 J	41.84 J	4.184 J	418.4 J
4.	The structure of water molecule is:	Angular	Linear	Trigonal	Trigonal pyramidal
5.	Which of the given molecules has zero dipole moment?	NH ₃	CHCl ₃	H ₂ O	BF ₃
6.	The wave number of the light emitted by a certain source is 2 × 10 ⁶ m ⁻¹ . The wavelength of this light will be:	500 nm	500 m	200 nm	5 × 10 ⁷ m
7.	The molecules of CO ₂ in dry ice form the:	Ionic crystals	Covalent crystals	Molecular crystals	Any type of crystals
8.	Splitting of spectral lines when atoms are subjected to strong electric field is called:	Zeeman effect	Stark effect	Photoelectric effect	Compton effect
9.	NH ₃ shows a maximum boiling point among the hydrides of group-V elements due to:	Very small size of nitrogen	Lone pair of electrons present on nitrogen	Enhanced electronegative character of nitrogen	Pyramidal structure of NH ₃
10.	Number of molecules in one dm ³ of water is close to:	$\frac{6.02}{22.4} \times 10^{23}$	$\frac{12.04}{22.4} \times 10^{23}$	$\frac{18}{22.4} \times 10^{23}$	$55.6 \times 6.02 \times 10^{23}$
11.	The deviation of a gas from ideal behaviour is maximum at:	-10°C and 5 atm	-10°C and 2 atm	100°C and 2 atm	0°C and 2 atm
12.	Which one of the given is not purified by sublimation?	Naphthalene	Silicon dioxide	I ₂	Benzoic acid
13.	Solvent extraction is an equilibrium process and it is controlled by:	Law of mass action	The amount of solvent used	Distribution law	The amount of the solute
14.	27g of Al will react completely with how much mass of O ₂ to produce Al ₂ O ₃ .	8g of oxygen	16g of oxygen	32g of oxygen	24g of oxygen
15.	Many elements have fractional atomic masses. This is because:	The mass of the atom is itself fractional	Atomic masses are average masses of isobars	Atomic masses are average masses of isotopes	Atomic masses are average masses of isotopes proportional to their relative abundance
16.	The rate of reaction:	Increases as the reaction proceeds	Decreases as the reaction proceeds	Remain the same as the reaction proceeds	May decrease or increase as the reaction proceeds
17.	Stronger the oxidizing agent, greater is the:	Oxidation potential	Reduction potential	Redox potential	E.M.F of cell

CANCELLED